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of the solute and solvent molecules

and on how strongly the latter interact with each other. This suggests two important limiting cases for reactions in solution. For water at room temperature, k_1 is typically $10^9 - 10^{10} \text{ dm}^{-3} \text{ mol}^{-1} \text{ s}^{-1}$ and k_2 is around $10^{-9} - 10^{-10} \text{ dm}^{-3}$

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17.5: Kinetics of Reactions in Solution
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solution in which $E_a > 20 \text{ kJ/mol}$ are
likely to fall into this category.

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Activation Controlled ($k_3 \gg k_2$): 3

Alternatively, if the activation energy of the A+B reaction dominates the kinetics, and the reaction is

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LibreTexts 4 CHAPTER 1 The Basics
of Reaction Kinetics for

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Chemical kinetics is the study of
chemical processes and rates of
reactions. This includes the analysis of

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conditions that affect speed of a chemical reaction, understanding reaction mechanisms and transition states, and forming mathematical models to predict and describe a chemical reaction. The rate of a chemical reaction usually has units of sec^{-1} , however, kinetics experiments

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